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CO₂ hydrogenation to methanol over Zr- and Ce-doped indium oxide / Salomone, F.; Sartoretti, E.; Ballauri, S.; Castellino, M.; Novara, C.; Giorgis, F.; Pirone, R.; Bensaid, S.. - In: CATALYSIS TODAY. - ISSN 0920-5861. - 423:(2023). [10.1016/j.cattod.2023.01.030]

Availability:

This version is available at: 11583/2979834 since: 2023-07-04T12:58:02Z

Publisher:

Elsevier

Published

DOI:10.1016/j.cattod.2023.01.030

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1 **CO₂ hydrogenation to methanol over Zr- and Ce-doped indium oxide**

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7 **Abstract**

8 In recent decades, climate change has become a major issue that needs to be addressed. Many efforts have
9 been made on the reduction of CO₂ emissions and its conversion in energy carriers and high value-added
10 products such as methane, methanol, dimethyl-ether, and hydrocarbons. The present study focuses on the
11 development of catalysts for hydrogenating CO₂ to methanol, which is a useful chemical and an alternative
12 liquid fuel. According to the literature, In₂O₃-based catalysts are particularly selective in the hydrogenation of
13 CO₂ to methanol, reducing the production of CO even at high space velocities compared to the more common
14 ternary catalysts such as Cu/ZnO/Al₂O₃ or Cu/ZnO/ZrO₂. Therefore, the effects of CeO₂ and ZrO₂ on In₂O₃-
15 based catalysts were investigated in the present study. The In_xCe_{100-x} and the In_xZr_{100-x} mixed oxides catalysts
16 were synthesized via gel-oxalate coprecipitation by varying the atomic ratios between the elements.
17 Subsequently, they were analysed with several characterisation techniques to rationalise the catalytic
18 performances that were obtained by testing the samples in a fixed bed reactor under different reaction
19 conditions. The addition of different amounts Ce or Zr modified the structure and morphology of the samples
20 and promoted the adsorption of CO₂ from 1.8 mmol_{CO₂}·g_{cat}⁻¹ up to 10.6 mmol_{CO₂}·g_{cat}⁻¹. ZrO₂ stabilises the
21 structure and the results suggests that the greater specific activity (168 mg_{CH₃OH}·g_{In₂O₃}⁻¹·h⁻¹ at 300 °C and 2.5
22 MPa of In₄₀Zr₆₀) could be ascribed to the electronic promotion of Zr. On the contrary, the addition of CeO₂ did
23 not reveal a beneficial effect on the activity. Concerning the stability, In₂O₃-ZrO₂ binary oxides seemed to be

24 affected mainly by sintering; whereas $\text{In}_2\text{O}_3\text{-CeO}_2$ were affected by at least three deactivating phenomena:
25 sintering, reduction of In_2O_3 to metallic indium and coking. Consequently, the deactivation rate of these binary
26 oxides increased from $1.04 \cdot 10^{-2} \text{ h}^{-1}$ of the In_{100} to $4.13 \cdot 10^{-2} \text{ h}^{-1}$ of the $\text{In}_{40}\text{Ce}_{60}$.

27 **Keywords**

28 Indium oxide, ceria, zirconia, methanol, CO_2 hydrogenation

29 **Highlights**

- 30 • CeO_2 and ZrO_2 modify the morphology and boost the CO_2 adsorption rate.
- 31 • The electronic promotion effect of Zr enhances the methanol synthesis from CO_2 .
- 32 • ZrO_2 stabilises the structure and mitigates the deactivation phenomena.
- 33 • CeO_2 increases the reducibility of the samples without stabilising the structure.
- 34 • $\text{In}_2\text{O}_3\text{-CeO}_2$ samples are affected by sintering, formation of metallic In and coking.

35 **1 Introduction**

36 In recent years, greenhouse gas emissions have become a serious problem due to the progressive increase in
37 the global average temperature. Therefore, the mitigation of the climate change and the reduction of CO_2
38 emissions are crucial challenges to be faced. Consequently, energy production technologies are progressively
39 changing and, in particular, renewable energy sources, such as wind or photovoltaic, are progressively
40 replacing non-renewable fossil energy sources. However, fossil fuels are still widely used in many industrial
41 fields and in the transport sector, where their full replacement is tricky. The goal of the society is the
42 achievement of the closed carbon cycle by producing alternative fuels from CO_2 , which can be captured from
43 flue gases or from the environment. Power-to-Fuels and Emission-to-Fuels processes aim to exploit the
44 electricity generated from renewable energy sources to produce an energy carrier such as H_2 [1–4], which can

45 be exploited in CO₂ hydrogenation processes to obtain alternative fuels or sustainable value-added chemicals
46 (e.g. methane, methanol, dimethyl-ether, liquid hydrocarbons, light olefins, etc.) [5–14].

47 Methanol is one of the most important components in the chemical industry, as it is widely used as a solvent
48 and chemical agent for the synthesis of other products [15–17]. The CO₂ hydrogenation to methanol is an
49 exothermic reaction and it is thermodynamically unfavoured by the temperature increase; instead, the reverse
50 water gas shift reaction is endothermic and thermodynamically favoured at higher temperatures. The overall
51 process involves three main reactions: the CO₂ and the CO hydrogenation to methanol and the reverse water
52 gas shift, as reported in equations (1), (2) and (3), respectively.



53 The CO₂ hydrogenation process is extensively studied in the literature and many articles support its technical
54 and economic feasibility for industrial scale production [5,8,16,18–21]. Cu-based catalysts (i.e. Cu/ZnO/Al₂O₃
55 or Cu/ZnO/ZrO₂) are the most studied for this application [17,22–24]. However, some researchers have lent
56 particular attention to In₂O₃-based catalysts [13,25,26].

57 More specifically, Martin *et al.* [26] have demonstrated the superior efficiency in CO₂ hydrogenation to
58 methanol of In₂O₃-based catalysts, achieving 100% methanol selectivity under industrial relevant reaction
59 conditions [25,26]. They have suggested that the reaction mechanism consists in the formation and annihilation
60 of oxygen vacancies on the surface of the catalytic material that enhance the catalytic performances and
61 stabilises In₂O₃ [26,27]. Moreover, they have confirmed that the indium oxide deposited on a zirconia support
62 showed excellent stability over a 1000-h endurance test under the reaction conditions [26]. Hence, according
63 to these recent studies, In₂O₃-ZrO₂ catalysts appear to be very promising materials for sustainable methanol
64 production. Moreover, the increase of CO₂ adsorption capacity of In₂O₃ supported on CeO₂ could be a
65 promising aspect for the capacity of converting CO₂ into methanol [27]. Other researchers have demonstrated
66 the beneficial effects of the addition of CeO₂ into the catalysts [28], and in the literature, it is known that the

67 high catalytic activity of ceria is often linked to its ability to acquire and release oxygen atoms forming oxygen
68 vacancies [29,30].

69 Therefore, this work aims to study the behaviour of co-precipitated $\text{In}_2\text{O}_3\text{-CeO}_2$ and $\text{In}_2\text{O}_3\text{-ZrO}_2$ binary oxide
70 catalysts for the methanol synthesis process to reduce CO production and maximize methanol productivity.
71 Furthermore, the goal is to rationalise the experimental evidence on both activity and stability of the samples
72 by means of different characterisation techniques.

73 **2 Methodology**

74 **2.1 Synthesis of the catalysts**

75 All the catalysts were prepared by gel-oxalate co-precipitation varying the atomic ratios of metals. A 0.1 M
76 solution of metal nitrates (i.e., $\text{In}(\text{NO}_3)_3 \cdot n\text{H}_2\text{O}$ and $\text{Ce}(\text{NO}_3)_3 \cdot 6\text{H}_2\text{O}$ or $\text{ZrO}(\text{NO}_3)_2 \cdot n\text{H}_2\text{O}$) in ethanol with the
77 proper atomic ratio and a 1 M solution of oxalic acid in ethanol were prepared [23,24,31]. After that, the oxalic
78 acid solution (an excess of 20 % of solution was used [32,33]) was rapidly poured ($\sim 1 \text{ mL} \cdot \text{s}^{-1}$) into the nitrate
79 solution, which was kept at room conditions under vigorous stirring, and the gel was aged for 2 h at these
80 conditions [32,33]. The precipitate was separated from the solvent by centrifugation and then it was dried at
81 $60 \text{ }^\circ\text{C}$ overnight. Finally, the dried precipitate was calcined at $360 \text{ }^\circ\text{C}$ for 4 h (heating rate: $10 \text{ }^\circ\text{C} \cdot \text{min}^{-1}$) and
82 then ground in an agate mortar for obtaining the catalyst powder. For the sake of clarity, the binary oxide
83 catalysts were labelled $\text{In}_x\text{Ce}_{100-x}$ and $\text{In}_x\text{Zr}_{100-x}$, where x represents the nominal atomic percentage of indium;
84 in this work x was set equal to 0, 40, 60, 80 and 100 at.%.

85 **2.2 Characterization of the catalysts**

86 Textural properties of calcined and spent catalysts were investigated by N_2 physisorption at 77 K using a
87 Micromeritics Tristar II ASAP 3020 analyser. Each sample was previously outgassed in a Micromeritics
88 FlowPrep 060 at $200 \text{ }^\circ\text{C}$ for 2 h flowing nitrogen to remove adsorbed contaminants and moisture from the
89 surface. The specific surface area was determined according to the Brunauer-Emmett-Teller theory [34,35]

90 and the porosimetry analysis was performed by applying the Barrett-Joyner-Halenda algorithm to the
91 desorption branch of the isotherm [36].

92 The X-ray diffractograms of the powders were collected at room conditions by using a Philips X'Pert PW3040
93 diffractometer operating at 40 kV and 40 mA, using a Ni β -filtered Cu-K α radiation characterized by a
94 wavelength (λ) equal to $1.5406 \cdot 10^{-10}$ m. The XRD data were collected over the 2θ angle range of 20° – 80° ,
95 with a scan step size of $0.013^\circ 2\theta$ and a time per step equal to 0.3 s, while the average crystallite size of each
96 phase was calculated according to Scherrer's equation [37] and the lattice strain was evaluated by using the
97 Williamson-Hall's equation [38,39].

98 The catalysts were observed with a field emission scanning electron microscope (FE-SEM) Zeiss Merlin
99 equipped with a Gemini-II column, for the purpose of analysing their morphology. Moreover, energy
100 dispersive X-ray spectroscopy (EDS) was employed to determine the elemental composition of portions of the
101 catalysts. In addition, the elemental composition of the catalysts was determined by using an inductively
102 coupled plasma mass spectrometer (iCAP Q ICP-MS, Thermo Fisher Scientific). Briefly, each calcined
103 catalyst (~100 mg) was digested in 30 mL of an acid aqueous solution (0.25 M L-ascorbic acid and 0.5 M
104 H₂SO₄) at room temperature overnight [40]. Each solution was diluted by using deionised water to have 100
105 mL of concentrated solution. After that, each concentrated solution was diluted using deionised water to
106 achieve the concentration of each element within the calibration range (i.e., 100 – 2000 ppb), then these
107 samples were analysed by using ICP-MS.

108 Temperature-programmed measurements were carried out using a Thermoquest TPD/R/O 1100 analyser
109 equipped with a thermal conductivity detector (TCD). Firstly, the reducibility properties of the samples were
110 investigated by means of H₂-TPR. Each catalyst (~20 mg of calcined sample) was placed in a quartz tube
111 reactor and pre-treated at 350°C ($+10^\circ\text{C}\cdot\text{min}^{-1}$) for 1 h in He flow ($20\text{ mL}\cdot\text{min}^{-1}$) to simulate the pre-treatment
112 prior to the catalytic tests. After that the sample was cooled to 40°C and the H₂-TPR measurement was
113 performed by using a constant flowrate of $20\text{ mL}\cdot\text{min}^{-1}$ of 5.00 vol.% H₂/Ar and by heating the oven from 40°C
114 $^\circ\text{C}$ to 900°C ($+10^\circ\text{C}\cdot\text{min}^{-1}$). Furthermore, the affinity of the CO₂ with the surface of the spent catalysts was
115 investigated by means of CO₂-TPD measurements. To perform the analysis, a mass of approximately 100 mg
116 was placed in the quartz tube reactor between two layers of quartz wool and cleaned with He ($20\text{ mL}\cdot\text{min}^{-1}$) at

117 300 °C (+10 °C·min⁻¹) for 1 h. Subsequently, the sample was cooled down to 40 °C in He and immediately
118 after it was saturated using pure CO₂ (20 mL·min⁻¹) at 40 °C for 1 h. Once the catalyst was saturated, it was
119 cleaned with He (20 mL·min⁻¹) at 40 °C for 1 h. Finally, the CO₂-TPD was performed by flowing 20 mL·min⁻¹
120 ¹ of He and heating the sample from 40 °C to 900 °C (+10 °C·min⁻¹).

121 The amount of carbonaceous compounds on the surface of the spent catalysts was estimated using temperature-
122 programmed combustion (TPC) analyses. Each sample (~50 mg) was placed in a U-shaped quartz tube and
123 the measurement was performed by flowing 100 NmL·min⁻¹ of 5 vol.% O₂/N₂ from room temperature to 750
124 °C (heating rate: +10 °C·min⁻¹). The gas at the outlet of the reactor was analysed by means of an ABB Uras 14
125 gas analyser to measure the concentration of CO and CO₂.

126 The surface of the catalysts was investigated via X-ray photoelectron spectroscopy (XPS) employing a PHI
127 5000 Versa Probe equipment, using, for survey scans, a band-pass energy of 187.85 eV, a take-off angle of
128 45° and an X-ray spot diameter of 100 µm; high resolution (HR) spectra were collected using a band-pass
129 energy of 23.50 eV.

130 Raman analyses of the In_xCe_{100-x} catalysts were carried out using a Renishaw InVia Reflex micro-Raman
131 spectrometer to study the defectiveness of the materials. The instrument is equipped with a 100 mW solid-state
132 laser emitting monochromatic light (wavelength: 514.5 nm). The Raman spectra were collected under a 5x
133 objective with a total acquisition time of 225 s. The results were elaborated with the Renishaw software WiRE
134 3.4. The D/F_{2g} ratio was calculated as previously described [41], by fitting the defect band with three
135 Lorentzian peaks located around 560, 590 and 620 cm⁻¹. An additional Lorentzian peak centred at 625 cm⁻¹
136 was introduced to consider a typical feature of In₂O₃ in this position; the area of the latter component, which
137 does not contribute to the D band area, was set proportional to the main In₂O₃ peak at 308 cm⁻¹. It is worth
138 noting that In_xZr_{100-x} samples could not be analysed due to excessive fluorescence.

139 The Fourier-transform infrared (FT-IR) analyses of the In_xZr_{100-x} samples were performed by means of a
140 Bruker Invenio S spectrophotometer equipped with a Mercury-Cadmium-Telluride (MCT) IR detector. Each
141 sample was mixed with KBr in a mass ratio of 1:4 to obtain a homogeneous powder. After that, some powder
142 was pelletised at 44 kN by producing a small thin tablet with an apparent density of 20 mg·cm⁻². The IR spectra
143 were always collected at room temperature between 4000 cm⁻¹ and 400 cm⁻¹ with a resolution of 2 cm⁻¹. Prior

144 to the measurements, each sample was outgassed at 0.01 Pa for 30 min at room temperature or at 100 °C, 200
145 °C or 300 °C.

146 **2.3 Catalytic tests**

147 The catalytic activity of each sample was investigated in a fixed-bed stainless steel reactor (catalyst load: 1.0
148 g). The catalytic bed had an annular section (i.d. 3 mm, o.d. 8 mm) due to the presence of an innertube, in
149 which a thermocouple was inserted for the measurement of the reaction temperature within the catalytic bed.

150 Concerning the catalytic tests, each sample was pre-treated in situ at 350 °C for 1 h by using pure nitrogen (30
151 NL·h⁻¹) at 0.2 MPa. Subsequently, each sample was tested continuatively (stability test) for 20 h at 275 °C, 2.5
152 MPa and 20 NL·g_{cat}⁻¹·h⁻¹ (inlet composition: 60 vol.% H₂, 20 vol.% CO₂ and 20 vol.% N₂). After the
153 stabilization, activity tests were performed on each catalyst at 2.5 MPa, 20 NL·g_{cat}⁻¹·h⁻¹ (inlet composition: 60
154 vol.% H₂, 20 vol.% CO₂ and 20 vol.% N₂) varying the reaction temperature between 200 °C and 300 °C. At
155 the end of the catalytic tests, the catalyst was removed from the reactor, sieved to remove quartz wool fibres
156 and collected to characterise it.

157 The gases at the outlet of the reactor were split in two streams. On the one side, a portion of the gases was
158 directly analysed by using a gas chromatograph (7890B GC System, Agilent Technologies) equipped with a
159 heated transfer line (120 °C, atmospheric pressure), a two-columns separation system (HP-PLOT/Q and HP-
160 PLOT Molesieve) connected to a thermal conductivity detector (TCD) and a flame ionization detector (FID).
161 On the other side, water and methanol were condensed in a small tank (room temperature and 2.5 MPa), then
162 the gaseous stream was completely dehydrated using a silica gel trap (room temperature and atmospheric
163 pressure), and it was eventually analysed with an in-line X-STREAM EMERSON gas analyser equipped with
164 two nondispersive infrared (NDIR) sensors and a thermal conductivity detector (TCD) for monitoring CO,
165 CO₂ and H₂ concentrations, respectively.

166 The CO₂ conversion (ζ_{CO_2} , dimensionless), the selectivity (σ_i , dimensionless) and the yield (η_i , dimensionless)
167 of a generic i-th product were evaluated by means of equations (4), (5) and (6), respectively. For the sake of
168 clarity, $\dot{n}_{CO_2,in}$ (mol·h⁻¹) and $\dot{n}_{CO_2,out}$ (mol·h⁻¹) are the inlet and outlet CO₂ molar flow rates, respectively; $N_{C,i}$

169 (dimensionless) is the number of carbon atoms of the i -th chemical species and $\dot{n}_{i,out}$ (mol·h⁻¹) is the outlet
 170 molar flow rate of the i -th compound. Normal conditions were assumed equal to 0 °C and 0.1 MPa.

$$\zeta_{CO_2} = \frac{\dot{n}_{CO_2,in} - \dot{n}_{CO_2,out}}{\dot{n}_{CO_2,in}} \quad (4)$$

$$\sigma_i = \frac{N_{C,i} \cdot \dot{n}_{i,out}}{\dot{n}_{CO_2,in} - \dot{n}_{CO_2,out}} \quad (5)$$

$$\eta_i = \zeta_{CO_2} \cdot \sigma_i = \frac{N_{C,i} \cdot \dot{n}_{i,out}}{\dot{n}_{CO_2,in}} \quad (6)$$

171 The molar fractions of CO₂, N₂, CO, CH₄, CH₃OH and CH₃OCH₃ were estimated by using calibration curves,
 172 the molar fraction of water was evaluated by means of the stoichiometry of the reactions and by taking into
 173 account the vapour pressure in the condensation tank; whilst the molar fraction of H₂ was evaluated by
 174 assuming that the sum of the molar fractions is equal to 1, because the sensitivity of the TCD detector to a
 175 variation of H₂ with respect to He (carrier) is too small due to the similar thermal conductivity of the two gases.
 176 The total outlet molar flow rate (\dot{n}_{out} , mol·h⁻¹) was estimated by assuming that the molar flow rate of the inert
 177 species (i.e., N₂) did not vary during the process, as reported in equation (7); where \dot{n}_{in} (mol·h⁻¹) is the total
 178 inlet molar flow rate, $y_{N_2,in}$ and $y_{N_2,out}$ are respectively the inlet and outlet N₂ molar fraction.

$$\dot{n}_{out} = \dot{n}_{in} \cdot \frac{y_{N_2,out}}{y_{N_2,in}} \quad (7)$$

179 Lastly, the uncertainty on the calculations was propagated according to equation (8), where $f(x_1, \dots, x_n)$ is a
 180 function of n variables (x_j), Δf is the uncertainty of the result of the function $f(x_1, \dots, x_n)$ and Δx_j is the
 181 uncertainty of a generic j -th variable.

$$\Delta f(x_1, \dots, x_n) = \left[\sum_{j=1}^n \left(\frac{\partial f}{\partial x_j} \cdot \Delta x_j \right)^2 \right]^{0.5} \quad (8)$$

182 2.4 Deactivation kinetic model

183 According to the literature [42], to study the deactivation phenomena, a 1st order kinetic of the reaction rate
 184 and a 1st order kinetic of the deactivation rate independent from the concentration of the reactants have been
 185 considered, as reported in the system of equations (9). In those equations, r_{CO_2} (mol_{CO₂}·m⁻³·h⁻¹) is the CO₂
 186 conversion rate, k' ((mol_{CO₂}·m⁻³)^{1- n} ·h⁻¹) is the Arrhenius' temperature-dependent kinetic constant of the

187 reaction rate, C_{CO_2} ($\text{mol}_{CO_2}\cdot\text{m}^{-3}$) is the CO_2 concentration in the gas phase, n (dimensionless) is the order of
 188 reaction, a (dimensionless) is the activity of the catalyst, da/dt (h^{-1}) is the deactivation rate, k_d ($(\text{mol}_{CO_2}\cdot\text{m}^{-3})^{-1}\cdot$
 189 h^{-1}) is the Arrhenius' temperature-dependent constant of deactivation, d (dimensionless) is the order of
 190 deactivation and m (dimensionless) measure the concentration dependency of the deactivation rate.

$$\begin{cases} r_{CO_2} = -k' \cdot C_{CO_2}^n \cdot a \\ \frac{da}{dt} = -k_d \cdot C_{CO_2}^m \cdot a^d \end{cases} \text{ where } n = 1, m = 0 \text{ and } d = 1 \quad (9)$$

191 By integrating the 2nd equation of the system of equations (9) from the initial time ($t = 0$ h) to a generic time t
 192 (h) of the time-on-stream, equation (10) could be obtained; where a_0 is the initial activity that was always
 193 assumed unitary.

$$a(t) = a_0 \cdot e^{-k_d \cdot t} \quad (10)$$

194 By considering the design of the reactor and the experimental conditions, the solid phase could be assumed
 195 stationary, while the fluid phase could be assumed as a plug flow with constant flow rate [42]. Therefore, the
 196 performance equation of the process could be expressed as reported in equation (11), where V_{cat} (m^3) is the
 197 volume of the catalytic bed in the reactor.

$$\frac{V_{cat}}{\dot{n}_{CO_2,in}} = \int_{t=0}^t \frac{d\zeta_{CO_2}}{-r_{CO_2}} \quad (11)$$

198 By substituting r_{CO_2} with the 1st equation of the system of equations (9) and substituting the activity with the
 199 equation (10), equation (11) could be integrated and rearranged to obtain a linearised expression as reported
 200 in equation (12) [42]; where τ' (h) is the average residence time of the gas. This latter equation could be used
 201 to estimate the deactivation constant rate of each catalyst during the time-on-stream.

$$\ln \left[\ln \left(\frac{1}{1 - \zeta_{CO_2}} \right) \right] = \ln(k' \cdot \tau') - k_d \cdot t \quad (12)$$

202 Lastly, the initial CO_2 conversion ($\zeta_{CO_2,0}$) obtained from this model could be estimated with equation (13) by
 203 assuming $t = 0$ h.

$$\zeta_{CO_2,0} = 1 - e^{-k' \cdot \tau'} \quad (13)$$

204 **2.5 Kinetic analysis**

205 To rationalise the results of the catalytic activity tests, the apparent activation energies for the CO₂
 206 hydrogenation reactions to CO and methanol could be obtained from the catalytic tests at different reaction
 207 temperatures (200 – 300 °C). According to the 1st expression of the system of equations (9) the kinetic factor
 208 could be expressed according to the Arrhenius' equation as reported in equation (14). In this equation, R_i
 209 ($\text{mol}_i \cdot \text{kg}_{\text{cat}}^{-1} \cdot \text{h}^{-1}$) is the reaction rate, ρ_{cat} ($\text{kg}_{\text{cat}} \cdot \text{m}^{-3}$) is the catalytic bed density, $k'_{\infty,i}$ ($(\text{mol}_i \cdot \text{m}^{-3})^{1-n} \cdot \text{h}^{-1}$) is the
 210 pre-exponential factor, $E_{A,i}$ ($\text{kJ} \cdot \text{mol}^{-1}$) is the apparent activation energy, C_i ($\text{mol}_i \cdot \text{m}^{-3}$) is the concentration of
 211 the i -th compound, n (dimensionless) is the order of reaction with respect to the i -th species and $k^*_{\infty,i}$ ($\text{mol}_i \cdot \text{kg}_{\text{cat}}^{-1}$
 212 $\cdot \text{h}^{-1}$) represents the apparent pre-exponential factor.

$$R_i = r_i \cdot \rho_{\text{cat}}^{-1} = k'_{\infty,i} \cdot \exp\left(-\frac{E_{A,i}}{R \cdot T}\right) \cdot C_i^n \cdot \rho_{\text{cat}}^{-1} = k^*_{\infty,i} \cdot \exp\left(-\frac{E_{A,i}}{R \cdot T}\right) \quad (14)$$

213 The equation (14) could be linearised to obtain equation (15). Thus, the apparent pre-exponential factors and
 214 the apparent activation energies for each catalyst could be estimated by fitting the overall experimental reaction
 215 rates of CO₂, CO and methanol on the Arrhenius' plot.

$$\ln(R_i) = \ln(k^*_{\infty,i}) - \frac{E_{A,i}}{R} \cdot \frac{1}{T} \quad (15)$$

216 It is well known that the pre-exponential factor and the apparent activation energy are strongly correlated
 217 parameters. Hence, if the reaction mechanism on the different catalysts does not change significantly, a unique
 218 value of the apparent activation energy for all the catalysts can be determined for each reaction to decouple the
 219 two parameters. The objective function of this latter optimization is presented in equation (16), where $E_{A,i}$
 220 ($\text{kJ} \cdot \text{mol}^{-1}$) is the apparent activation anergy of the i -th reaction (for the sake of clarity, i could be CO₂, CO or
 221 CH₃OH), $k^*_{\infty,i,j}$ ($\text{mol}_i \cdot \text{kg}_{\text{cat}}^{-1} \cdot \text{h}^{-1}$) is the apparent pre-exponential factor of the i -th reaction for the j -th catalyst,
 222 $T_{j,l}$ (K) is the l -th reaction temperature for the j -th catalyst and $R_{i,j,l,sp}$ ($\text{mol}_i \cdot \text{kg}_{\text{cat}}^{-1} \cdot \text{h}^{-1}$) is the experimental
 223 reaction rate of the i -th species for the j -th catalyst at the l -th reaction temperature.

$$f(E_{A,i}, k^*_{\infty,i,1}, \dots, k^*_{\infty,i,M}) = \min \sum_{j=1}^M \sum_{l=1}^L \left\{ \ln(R_{i,j,l,sp}) - \left[\ln(k^*_{\infty,i,j}) - \frac{E_{A,i}}{R} \cdot \frac{1}{T_{j,l}} \right] \right\}^2 \quad (16)$$

224 3 Results and Discussion

225 3.1 Physico-chemical characterization of the catalysts

226 All the catalysts were analysed with several characterization techniques to measure and evaluate different
227 physico-chemical aspects that have been correlated to the synthesis technique and to the catalytic activity and
228 stability. First of all, the elemental composition of the calcined samples was measured with ICP-MS, since the
229 degree of hydration of two precursors was unknown, as reported previously in the methodology; for the sake
230 of clarity, those nitrates were assumed trihydrate. The results in terms of the elemental composition of the
231 calcined catalysts are detailed in Table S1, while Table 1 summarises the indium atomic content in the binary
232 oxide samples. It is worth noting that the actual elemental composition of the $\text{In}_x\text{Ce}_{100-x}$ samples is close to the
233 nominal one, while the elemental composition of the $\text{In}_x\text{Zr}_{100-x}$ samples differs uniformly from the nominal
234 values. Hence, the hydration degree of the zirconyl nitrate was probably greater than the assumed value.

235

236 Table 1. Indium atomic content in the binary oxide calcined samples determined by means of ICP-MS
237 measurements.

In/(In+M ^a) atomic ratio	Binary oxide catalysts					
	$\text{In}_{40}\text{Ce}_{60}$	$\text{In}_{60}\text{Ce}_{40}$	$\text{In}_{80}\text{Ce}_{20}$	$\text{In}_{40}\text{Zr}_{60}$	$\text{In}_{60}\text{Zr}_{40}$	$\text{In}_{80}\text{Zr}_{20}$
Nominal	0.40	0.60	0.80	0.40	0.60	0.80
Actual	0.39 ± 0.02	0.60 ± 0.02	0.85 ± 0.01	0.49 ± 0.01	0.71 ± 0.01	0.87 ± 0.01

^a M represents respectively Ce or Zr in the $\text{In}_x\text{Ce}_{100-x}$ or $\text{In}_x\text{Zr}_{100-x}$ mixed catalysts.

238

239 The textural properties of both calcined and spent samples were evaluated by means of N_2 physisorption
240 measurements; the results are summarised in Table 2. As far as the $\text{In}_x\text{Ce}_{100-x}$ catalysts are concerned, the
241 specific surface area (S_{BET} , $\text{m}^2\cdot\text{g}_{\text{cat}}^{-1}$) and the average pore size (d_p , nm) of the calcined samples did not show a
242 clear trend, while the total pore volume (V_{BJH} , $\text{cm}^3\cdot\text{g}_{\text{cat}}^{-1}$) slightly increases as the indium content rises. After
243 the catalytic tests, the specific surface area of those samples approximately halved, while the mean pore size
244 approximately doubled. On the contrary, both calcined and spent $\text{In}_x\text{Zr}_{100-x}$ samples exhibited a high specific
245 surface area, showing a greater stability under reaction condition by increasing the zirconium content. It was

246 also observed that the average pore size of those catalysts did not change significantly, while the total pore
 247 volume increased slightly during the catalytic tests. Overall, these results indicate that zirconium reduces
 248 sintering phenomena and stabilises the structure of the binary oxides better than cerium both during the
 249 preparation of the samples and during the catalytic tests. The latter is an important feature to keep high catalytic
 250 performances under reaction conditions.

251

252 Table 2. Textural properties of calcined and spent catalysts estimated by means of N₂ physisorption.

Catalyst	Calcined samples			Spent samples		
	S _{BET} (m ² ·g _{cat} ⁻¹)	V _{BJH} (cm ³ ·g _{cat} ⁻¹)	d _p (nm)	S _{BET} (m ² ·g _{cat} ⁻¹)	V _{BJH} (cm ³ ·g _{cat} ⁻¹)	d _p (nm)
Ce ₁₀₀	59.6 ± 0.4	0.121	4.6	-	-	-
In ₄₀ Ce ₆₀	41.3 ± 0.3	0.132	10.1	19.1 ± 0.2	0.122	21.8
In ₆₀ Ce ₄₀	61.3 ± 0.3	0.159	8.2	26.7 ± 0.2	0.132	15.4
In ₈₀ Ce ₂₀	55.6 ± 0.3	0.167	9.8	21.5 ± 0.2	0.132	20.3
In ₁₀₀	35.6 ± 0.3	0.135	12.7	18.3 ± 0.1	0.137	29.4
In ₈₀ Zr ₂₀	68.8 ± 0.2	0.160	7.2	51.1 ± 0.3	0.165	10.3
In ₆₀ Zr ₄₀	63.8 ± 0.3	0.122	6.0	59.9 ± 0.4	0.124	6.2
In ₄₀ Zr ₆₀	78.4 ± 0.2	0.129	5.2	68.8 ± 0.2	0.132	6.0
Zr ₁₀₀	80.7 ± 0.3	0.091	3.4	-	-	-

253

254 The structural properties of the catalysts were investigated by using XRD. The X-ray diffractograms of the
 255 calcined and spent samples are shown in Figure S1 and in Figure 1, respectively. Concerning In_xCe_{100-x}
 256 samples, the diffraction peaks of the cubic structure of CeO₂ (PFD 01-081-0792) are clearly observable at
 257 28.542°, 33.075°, 47.475° and 56.332°. Besides, the diffraction peaks of the cubic structure of In₂O₃ (PFD 03-
 258 065 -3170) are located at 21.448°, 30.515°, 35.380°, 50.901° and 60.517°. On the contrary, ZrO₂ is present as
 259 amorphous and non-crystalline form in the In_xZr_{100-x} samples: in fact, they exhibited two large halos located
 260 approximately between 25° and 35° and between 45° and 60°.

261

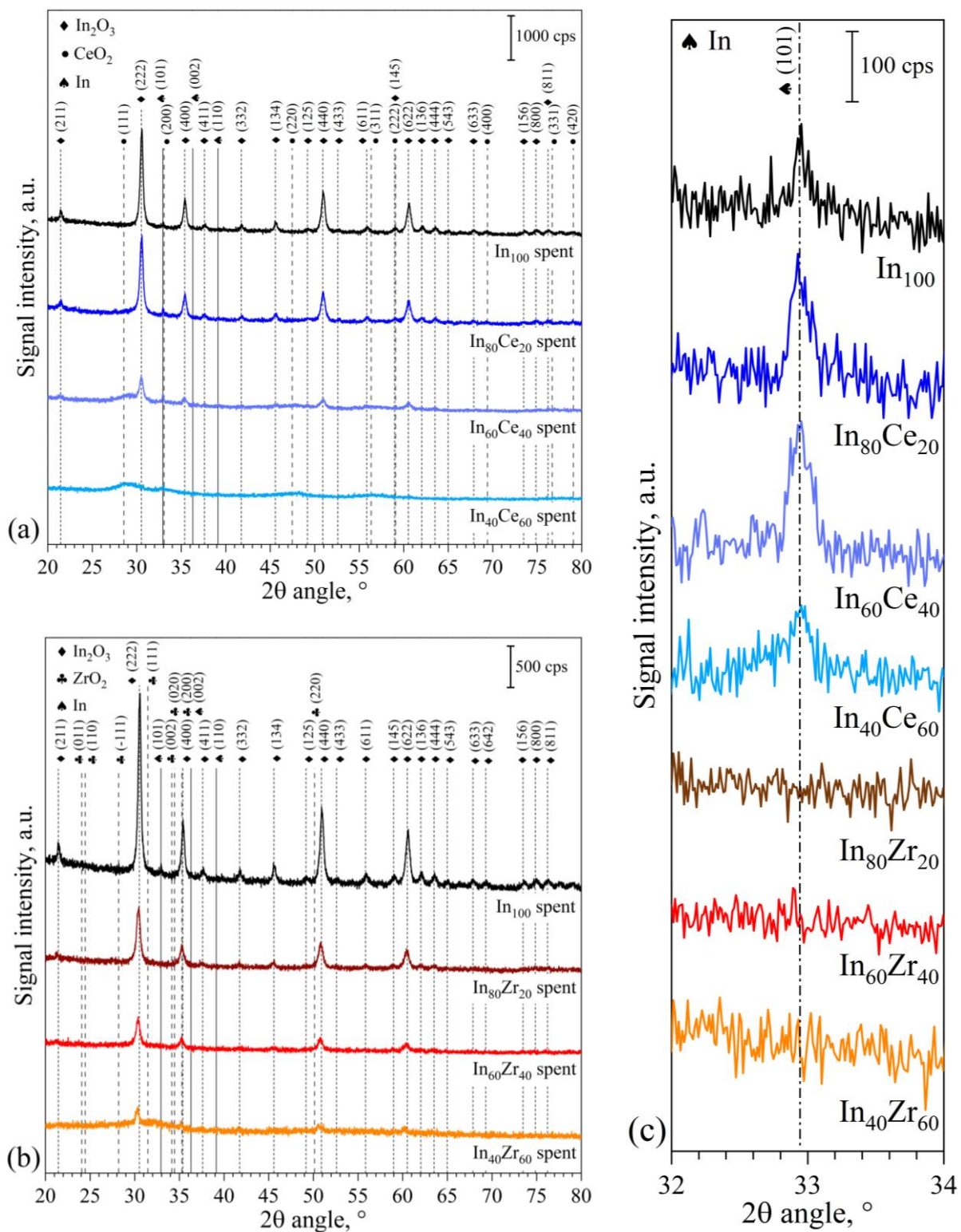


Figure 1. X-ray diffractograms of the spent (a) $\text{In}_x\text{Ce}_{100-x}$ and (b) $\text{In}_x\text{Zr}_{100-x}$ catalysts and (c) magnification of the XRD patterns of the spent samples between 32° and 34° of the 2θ angle.

263 The complete disappearance of the In_2O_3 peaks in the diffractogram of the $\text{In}_{40}\text{Zr}_{60}$ catalyst suggests that indium
264 may be incorporated in the amorphous structure of zirconia, forming a solid solution of $\text{In}_2\text{O}_3\text{-ZrO}_2$ (see Figure
265 S1). Metallic indium, Ce_2O_3 or sub-stoichiometric forms of ZrO_x were not detected by means of XRD in the
266 calcined samples. A closer inspection of the diffractograms of the spent catalysts reported in Figure 1 reveals
267 that the peaks of In_2O_3 and CeO_2 are more intense and narrower than the corresponding peaks in the calcined
268 samples, suggesting that sintering phenomena occurred under catalytic tests. On the other hand, ZrO_2 remained
269 amorphous without exhibiting clear detectable diffraction peaks. It is worth mentioning that the diffraction
270 peaks of the tetragonal structure of the metallic indium (PFD 03-065-9682), located at 32.930° , 36.301° and
271 39.131° , were observed in the XRD patterns of the spent $\text{In}_x\text{Ce}_{100-x}$ samples (see Figure 1a), while they were
272 not detected in the diffractograms of the spent $\text{In}_x\text{Zr}_{100-x}$ samples (see Figure 1b). In more detail, Figure 1c
273 displays a magnification of the XRD patterns of the aged samples between 32° and 34° , showing the peak
274 attributed to the Miller's indexes (101) of the metallic indium. This is a remarkable outcome that is related to
275 the activity and stability of the binary oxides as described in more detail in the following paragraphs and in
276 section 3.2.

277 As far as the crystalline structure is concerned, Table S2 summarises the average crystallite size, the lattice
278 strain and the shift of the most intense peak of diffraction of each oxide phase. More specifically, what is
279 interesting in this table is the growth of the crystallite size of In_2O_3 from ~ 3 nm to ~ 23 nm as the indium content
280 in the calcined samples increases. Moreover, the average crystallite size of In_2O_3 rose by approximately 30%
281 during the catalytic tests. Similarly, the average crystallite size of metallic indium in the spent $\text{In}_x\text{Ce}_{100-x}$
282 samples grew from 140 nm to 420 nm as the indium content increases. Regarding the size of the CeO_2
283 crystallites in the $\text{In}_x\text{Ce}_{100-x}$ samples, it remained steady at around 3 nm as the cerium content in the samples
284 varied. Lastly, it is worth mentioning that the peaks of CeO_2 , and in particular the more intense diffraction
285 peak located at 28.542° , shift proportionally towards larger angles as the indium content in the sample rises.
286 This positive shift has been ascribed to the formation of In-doped CeO_2 in which the indium atoms have an
287 atomic radius smaller than the atomic radius of cerium [43]; indeed, the consequent shrinkage of ceria fluorite
288 structure is responsible for the shift of XRD peaks [37]. On the contrary, Ce-doped In_2O_3 does not appear to
289 be formed as the diffraction peaks of In_2O_3 do not shift.

290 The XRD outcomes were confirmed by Raman spectroscopy. In the Raman spectra of the mixed oxides (see
 291 Figure 2), the intense F_{2g} feature of CeO_2 , associated with Ce-O symmetric stretching, is always present, while
 292 the typical vibrational modes of In_2O_3 were detected only in the samples presenting a high indium content [44].
 293 The remarkable broadening of CeO_2 F_{2g} peak and its shift from 464 cm^{-1} to 454 cm^{-1} in the spectra of the
 294 In_xCe_{100-x} mixed oxides are likely due to the incorporation of indium ions into ceria lattice, confirming the
 295 formation of an In-doped CeO_2 phase. Consistently, a significant increase of CeO_2 defect band was observed
 296 in the presence of In, as also proved by the increase of the D/F_{2g} area ratio. This parameter, calculated as the
 297 ratio between the area of the broad defect-induced band around 600 cm^{-1} and the area of the main F_{2g} peak, is
 298 widely employed to estimate defect abundance in ceria-based samples. An attempt was also made to semi-
 299 quantitatively evaluate the defect abundance in the In_2O_3 structure [29]. Since Gan et al. [45] correlated the
 300 intensity of the Raman feature at 367 cm^{-1} to the presence of oxygen vacancies in indium oxide, we calculated
 301 the ratio between the areas of this component and the one of the main In_2O_3 peak at 308 cm^{-1} ; due to the too
 302 low intensity of these peaks in the Ce-rich materials, this calculation was only possible for the In_{100} and $In_{80}Ce_{20}$
 303 samples. The area ratio increases upon Ce addition and after the catalytic tests, suggesting that more oxygen
 304 defects form in the In_2O_3 structure.

305

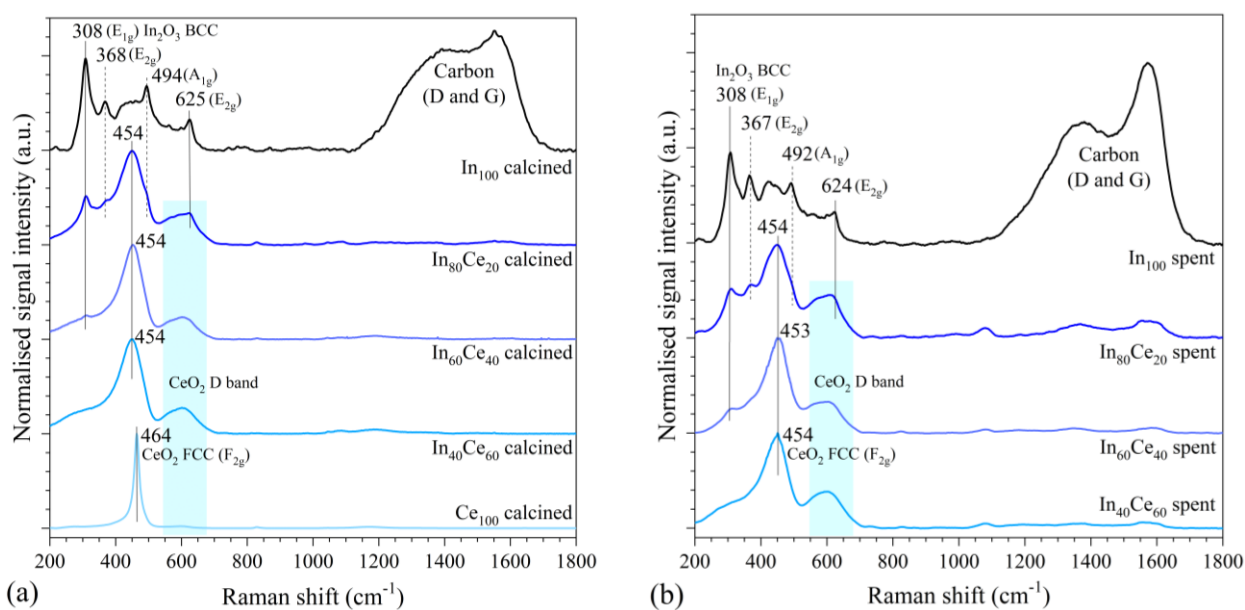


Figure 2. Raman spectra of (a) calcined and (b) spent In_xCe_{100-x} samples. For the sake of clarity, the signals were normalised with respect to the F_{2g} peak, except for the In_{100} , where the E_{1g} peak was used [44].

306

307 As aforementioned, metallic indium was detected only in the spent $\text{In}_x\text{Ce}_{100-x}$ samples by means of XRD
308 measurements, as can be clearly seen in Figure 1c. Moreover, what stands out in Figure 3 is that the H_2 -TPR
309 measurements confirmed the increased reducibility of those catalysts in a range of temperature that is
310 consistent with the reaction conditions. More specifically, the total H_2 consumption for the $\text{In}_x\text{Ce}_{100-x}$ samples
311 rises linearly from $1.06 \text{ mmol}_{\text{H}_2} \cdot \text{g}_{\text{cat}}^{-1}$ to $11.1 \text{ mmol}_{\text{H}_2} \cdot \text{g}_{\text{cat}}^{-1}$ as the indium content grows from 0 at.% to 100
312 at.%. In Figure 3b, the low-temperature peaks at $268 \text{ }^\circ\text{C}$, $292 \text{ }^\circ\text{C}$ and $344 \text{ }^\circ\text{C}$ correspond respectively to the
313 1.7%, 6.8% and 15.1% of the total H_2 consumption of each sample. This means that the amount of reduced
314 In_2O_3 at low temperature increases by reducing its content in the catalysts, causing a loss of active phase under
315 reaction conditions as discussed in the next section. Concerning the $\text{In}_x\text{Zr}_{100-x}$ binary oxides, their total H_2
316 consumption (i.e., $9.3 \text{ mmol}_{\text{H}_2} \cdot \text{g}_{\text{cat}}^{-1}$ for $\text{In}_{40}\text{Zr}_{60}$, $15.2 \text{ mmol}_{\text{H}_2} \cdot \text{g}_{\text{cat}}^{-1}$ for $\text{In}_{60}\text{Zr}_{40}$ and $17.3 \text{ mmol}_{\text{H}_2} \cdot \text{g}_{\text{cat}}^{-1}$ for
317 $\text{In}_{80}\text{Zr}_{20}$) is greater than the one of the $\text{In}_x\text{Ce}_{100-x}$ binary oxides, suggesting that the addition of zirconium also
318 increases the reducibility of In_2O_3 ; however, it seems to stabilise the structure because the reduction of In_2O_3
319 occurs at temperatures higher than the reaction temperature of the catalytic tests. Accordingly, as presented in
320 Figure 1c, no metallic indium was detected in the spent $\text{In}_x\text{Zr}_{100-x}$ samples. This is an important aspect because
321 the formation of metallic indium in the $\text{In}_x\text{Ce}_{100-x}$ binary oxides could be directly related to the progressive
322 deactivation of the samples as described more specifically in section 3.2.

323

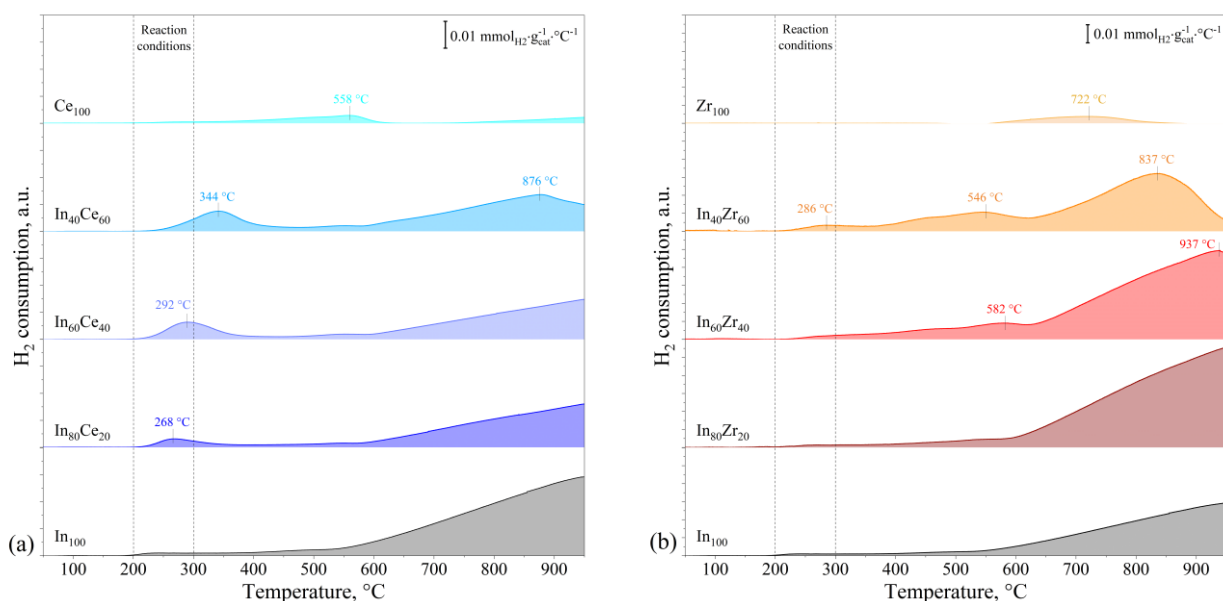


Figure 3. H₂-TPR of the calcined samples: (b) In_xCe_{100-x} and (c) In_xZr_{100-x}.

324

325 Moving on now to consider the results of the CO₂-TPD, Figure 4 displays the CO₂-TPD profiles of the spent
 326 samples. It is apparent from these profiles that the strength of the basic sites on the surface of the catalytic
 327 material changes completely in the In_xZr_{100-x} samples. The amount of CO₂ desorbed at high temperature (T >
 328 700 °C) increases from 0.29 mmol_{CO₂}/g_{cat} to 2.55 mmol_{CO₂}/g_{cat} by growing the ZrO₂ content in the catalysts;
 329 however, at these high temperatures the signal may be disturbed by the release of oxygen from the surface,
 330 that was not observed in the other samples. On the other hand, the addition of CeO₂ in the samples slightly
 331 reduces the strength of the CO₂ bonds with the superficial sites. The different desorption temperature could be
 332 related to the different chemisorbed species such as bicarbonate, bidentate- or tridentate-like carbonate species
 333 [46]. For the sake of completeness, Figure S2 shows the CO₂-TPD measurements of the calcined samples. It
 334 is worth noting that all the samples desorbed CO₂ at low temperature, but this feature changes in the spent
 335 catalysts probably due to the prolonged exposition to reaction conditions. Both calcined and spent samples
 336 desorbed the main amount of CO₂ between 400 °C and 600 °C.

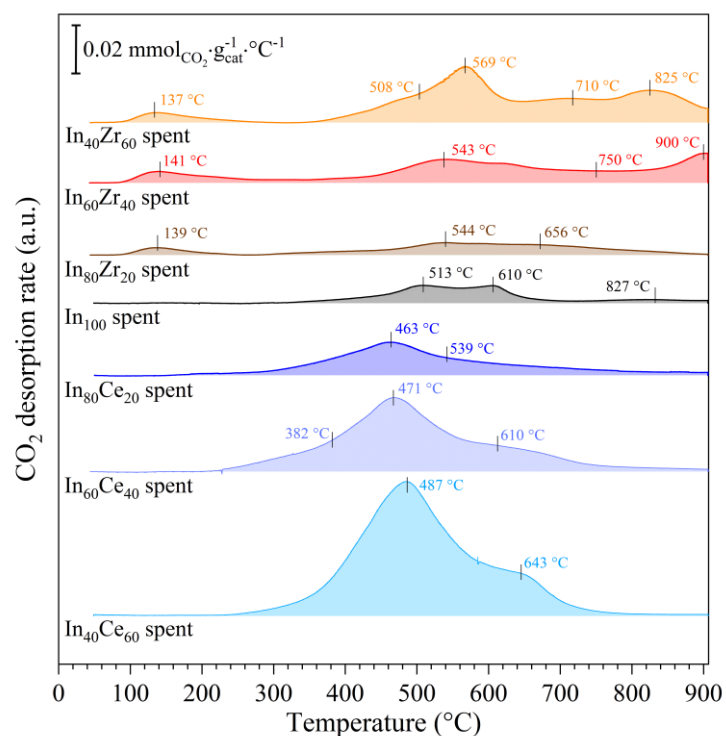


Figure 4. CO₂-TPD of the spent In_xCe_{100-x} and In_xZr_{100-x} catalysts.

337

338 Furthermore, as shown in Figure 5a, the desorbed amount of CO₂ increases proportionally as the indium
 339 content diminishes for both the binary oxides groups. In support of this result, XPS measurements were
 340 performed on the spent samples (see Figure S3, Figure S4, Figure S5 and Figure S6). As shown in Figure S3
 341 and Figure S4, the XPS HR spectra for the O1s and the C1s signals exhibited respectively a peak located at
 342 ~531.7 eV and a peak located at ~289 eV that grew by reducing the indium content in the samples. These
 343 signals have been ascribed to the presence of carbonates and to the organic C=O bond [47]. In fact, as shown
 344 in Figure 5b, the amount of CO₂ desorbed seems to be related to the fraction of the deconvoluted peak of the
 345 O1s signal located at ~531.7 eV. Moreover, as reported in the literature, a peak of the O1s signal located at
 346 ~532.5 eV could be related to the presence of hydroxyl species [46,48]; whereas a peak located at ~530.5 eV
 347 was attributed to O defects [26,46,48]. The presence of both hydroxyl groups and carbonates was confirmed
 348 by means of FT-IR measurements on the In_xZr_{100-x} samples as illustrated in Figure S7 and Figure S8. A closer
 349 inspection of the spectra revealed that the wide band between 3700 cm⁻¹ and 2900 cm⁻¹ is commonly attributed
 350 to the stretching vibrations of the hydroxyl groups of adsorbed water molecules [49]. Besides, the peaks located

351 at $\sim 1545\text{ cm}^{-1}$ and $\sim 1370\text{ cm}^{-1}$ have been ascribed to the presence of carbonates on the surface [50,51]. Lastly,
 352 the narrow peak located at 2337 cm^{-1} may be assigned to gaseous CO_2 trapped in closed channels [51] of the
 353 samples or to stretching vibrations of the Zr-OH bonds [52,53].
 354

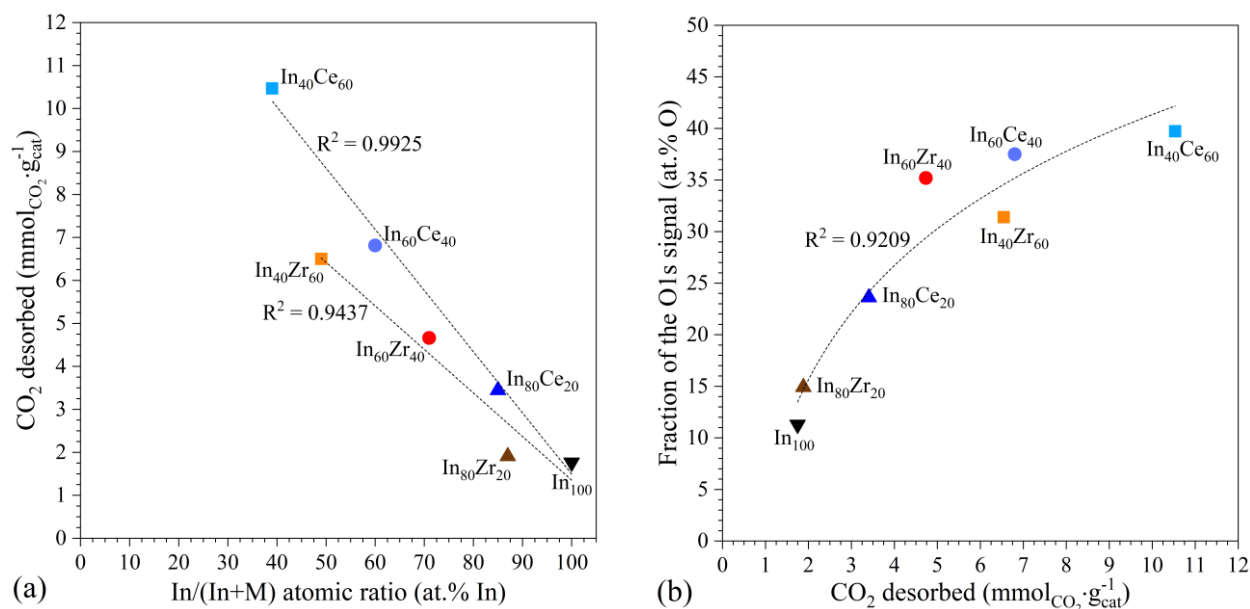


Figure 5. (a) Relationship between the indium content estimated by using ICP-MS (M represents Ce or Zr, see Table 1) and the CO₂ desorbed during the CO₂-TPD of the spent catalysts. (b) Relationship between the desorbed amount of CO₂ during the CO₂-TPD and the fraction of the O1s signal located at $\sim 531.7\text{ eV}$ obtained from the deconvolution of the XPS HR spectra (see Figure S3).

355

356 Further XPS HR spectra were collected for the In3d, Ce3d and Zr3d peaks of the spent catalysts and the results
 357 are shown in Figure S5, Figure S6a and Figure S6b, respectively. Concerning the In3d signals of the In_xCe_{100-x}
 358 samples, it is evident that there are no differences at the variation of the indium content. On the other hand,
 359 a small shift from 444 eV to 444.6 eV of the In3d_{5/2} of the In_xZr_{100-x} samples binding energy occurred by
 360 reducing the indium content; according to the literature, it may be related to the formation of chemisorbed
 361 species that participate into the overall CO₂ hydrogenation reaction pathway. Loh and Kerani (2019) have
 362 suggested that a possible reaction pathway on the surface of In₂O_{3-x}OH_y may involve the filling of an oxygen
 363 vacancy to form a bidentate-like carbonate with a C-O bond that may be activated by a neighbouring hydride
 364 species to form a formate group [46]. Moreover, the formation of In-C bonds in the spent In_xZr_{100-x} samples
 365 was detected by means of XPS, suggesting that the formation of oxygen defects (or vacancies) could be filled

366 by CO₂ forming a bidentate-like carbonate [26,46]. Lastly, the XPS HR spectra for the Ce3d and the Zr3d
367 peaks of the spent binary oxides were deconvoluted to estimate the average oxidation state and the results will
368 be further discussed in section 3.3.

369 Before proceeding to examine the catalytic performances of the binary oxide catalysts, another significant
370 aspect is the morphology of the samples. The different morphological nature of the two groups of catalysts is
371 clearly exemplified in Figure 6. The spent In₁₀₀ catalyst is made up of agglomerates of particles and it is evident
372 that the structure of the material is not completely different from the calcined In₁₀₀ sample (see Figure S9a).
373 The size of the particles of the spent In₁₀₀ sample ranges from ~15 nm to ~40 nm, which is consistent with the
374 average crystallite size estimated by means of XRD measurements (see Table S2). As far as the In_xCe_{100-x}
375 samples are concerned, they showed a mesoporous sponge-like structure, that thickens by increasing the Ce
376 content. The very small round-shaped particles have been assigned to CeO₂, while the more regular and larger
377 crystals have been attributed to the cubic structure of In₂O₃. On the other hand, In_xZr_{100-x} samples have a
378 completely different structure. Concerning the calcined samples, In₈₀Zr₂₀ and In₆₀Zr₄₀ showed a lamellar-like
379 structure, while In₄₀Zr₆₀ is characterised by intertwined fibres on which aggregates of small particles are
380 dispersed, as illustrated in Figure S9. The spent catalysts are made of a fibrous structure decorated with In₂O₃
381 crystals (see Figure 6e-g). The morphology of those samples changes completely under reaction conditions.

382

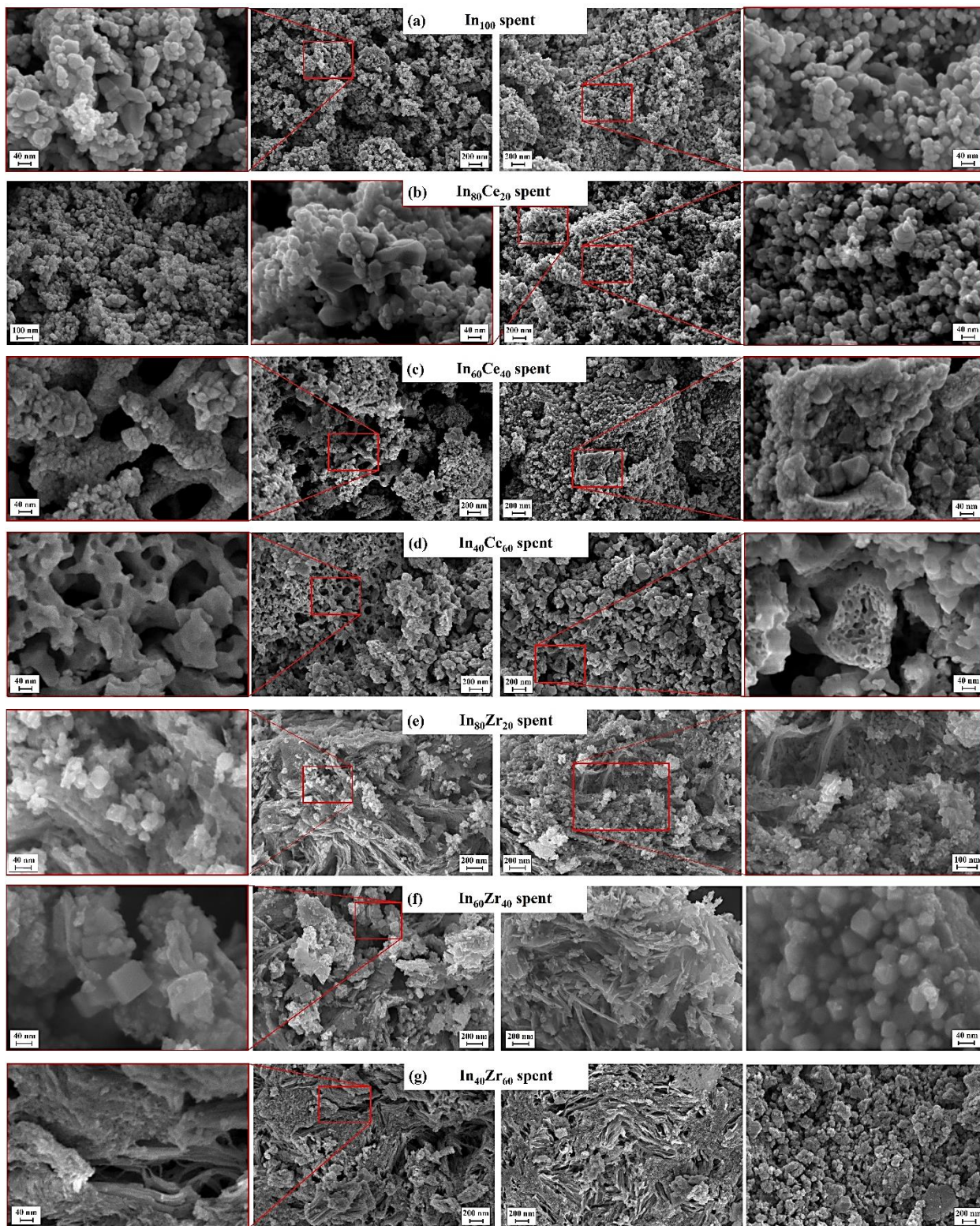


Figure 6. FE-SEM images of the spent catalysts: (a) In_{100} , (b) $\text{In}_{80}\text{Ce}_{20}$, (c) $\text{In}_{60}\text{Ce}_{40}$, (d) $\text{In}_{40}\text{Ce}_{60}$, (e) $\text{In}_{80}\text{Zr}_{20}$, (f) $\text{In}_{60}\text{Zr}_{40}$ and (g) $\text{In}_{40}\text{Zr}_{60}$.

384 3.2 *Catalytic stability*

385 So far, this paper has focused on the results of the characterisation techniques. The following section will
386 present the results of the stability tests performed on the binary oxide catalysts, attempting to provide an
387 explanation to the behaviour of the catalytic performances. Before proceeding to examine the stability tests, it
388 is important to restate that, prior to the tests, each catalyst was pre-treated in pure nitrogen at 300 °C for 1 h,
389 then it was subjected to the 20-h stability test at 275 °C, 2.5 MPa, and 20 NL·g_{cat}⁻¹·h⁻¹ (inlet gas composition:
390 60 vol.% H₂, 20 vol.% CO₂ and 20 vol.% N₂). The experimental results are shown in Figure 7. The In_xZr_{100-x}
391 catalysts exhibited a CO₂ conversion rate higher than pure In₂O₃, showing that the addition of zirconium
392 promotes the CO₂ hydrogenation reactions. On the contrary, In_xCe_{100-x} binary oxides exhibited a fall of the CO₂
393 conversion rate caused by the addition of cerium in the samples. As shown in Figure 7b, the In_xCe_{100-x} catalysts
394 exhibited a methanol selectivity greater than the In_xZr_{100-x} catalysts, but the methanol yield of the In_xZr_{100-x}
395 catalysts remained superior. Overall, the activity of the samples will be discussed in more detail in the next
396 section; so, focusing now on the stability of the samples, Figure 7a reveals a progressive loss of activity under
397 reaction conditions. The deactivation is caused by several phenomena, more precisely, it may be ascribed to
398 those listed below:

- 399 • a loss of the active surface area due to the shrinkage and the rearrangement of the structure,
- 400 • sintering of the particles that leads to a reduction of the dispersion of the active phase,
- 401 • the formation of metallic indium that results in a gradual decline of the active surface area,
- 402 • coking of the active surface area caused by the deposition of carbon species.

403

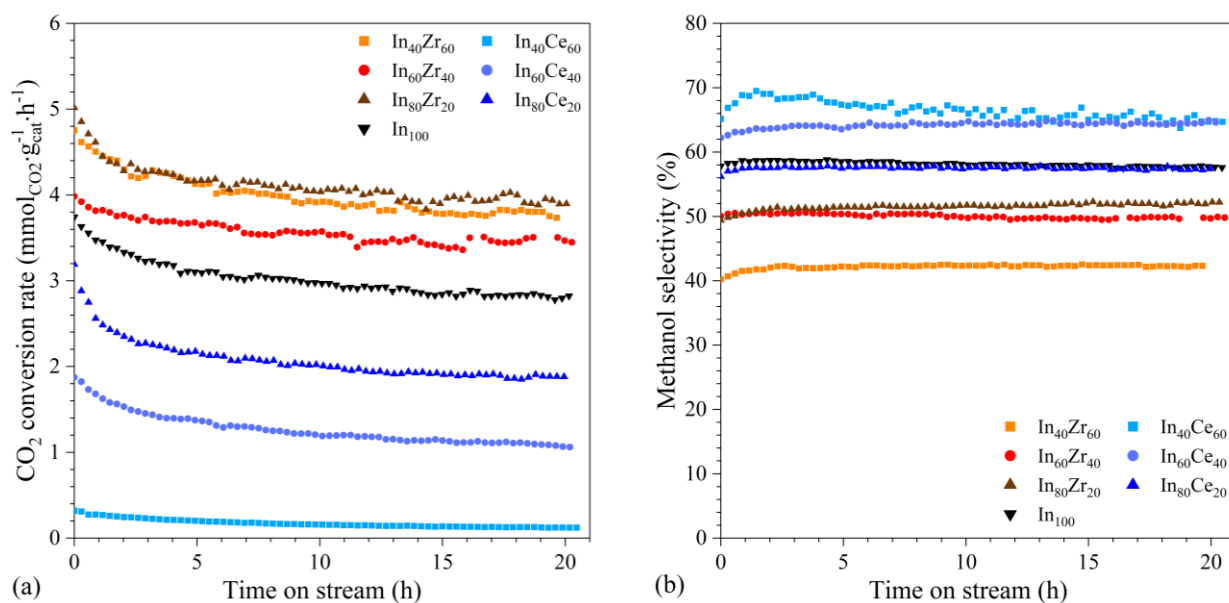


Figure 7. (a) CO₂ conversion rate and (b) methanol selectivity of the In_xCe_{100-x} and In_xZr_{100-x} catalysts during the 20-h stability test. Reaction conditions are 275 °C, 2.5 MPa and 20 NL·g_{cat}⁻¹·h⁻¹ (inlet gas composition: 60 vol.% H₂, 20 vol.% CO₂ and 20 vol.% N₂).

404

405 For the purpose of rationalising the causes of the deactivation, the CO₂ conversion profiles were fitted by using
 406 the equation (11) and the deactivation parameters are summarised in Table 3. What stands out from the table
 407 is the growth of the deactivation constant by increasing the cerium content in the samples. These results further
 408 support the idea that the addition of CeO₂ reduces the stability of the samples, while ZrO₂ stabilises the binary
 409 oxide structure and increases the overall activity in CO₂ hydrogenation.

410

411

Table 3. Deactivation parameters of the 1st order deactivation rate at 275 °C.

Catalyst	Deactivation constant (k_d), h ⁻¹	$\ln(k' \cdot \tau)$, dimensionless	Correlation coefficient (R^2)
In ₄₀ Ce ₆₀	$(4.13 \pm 0.13) \cdot 10^{-2}$	-6.553 ± 0.016	0.9308
In ₆₀ Ce ₄₀	$(2.14 \pm 0.10) \cdot 10^{-2}$	-4.736 ± 0.012	0.8713
In ₈₀ Ce ₂₀	$(1.60 \pm 0.11) \cdot 10^{-2}$	-4.293 ± 0.013	0.7476
In ₁₀₀	$(1.04 \pm 0.06) \cdot 10^{-2}$	-3.963 ± 0.007	0.8361
In ₈₀ Zr ₂₀	$(7.41 \pm 0.56) \cdot 10^{-3}$	-3.684 ± 0.007	0.7168
In ₆₀ Zr ₄₀	$(5.48 \pm 0.36) \cdot 10^{-3}$	-3.846 ± 0.004	0.7657
In ₄₀ Zr ₆₀	$(9.28 \pm 0.48) \cdot 10^{-3}$	-3.694 ± 0.006	0.8474

412

413 Moreover, the CO₂ conversion profiles showed a steep drop during the initial part of the tests, that might be
 414 caused by the reduction of the specific surface area and by the sintering of the particles. Hence, a reasonable
 415 approach to tackle this issue could be to find a relationship between the drop of the activity during the initial
 416 part of the stability test and the reduction of specific surface area measured by using N₂ physisorption. More
 417 precisely, as illustrated in Figure 8, the reduction of the specific surface area of all the samples (see Table 2)
 418 seems to be correlated to the sharp fall of CO₂ conversion during the stability tests. This finding could provide
 419 an explanation to the behaviour of the catalysts during the initial part of the stability test; however, they reveal
 420 that a progressive decline of the CO₂ conversion rate during the time on stream may be attributed to other
 421 causes.
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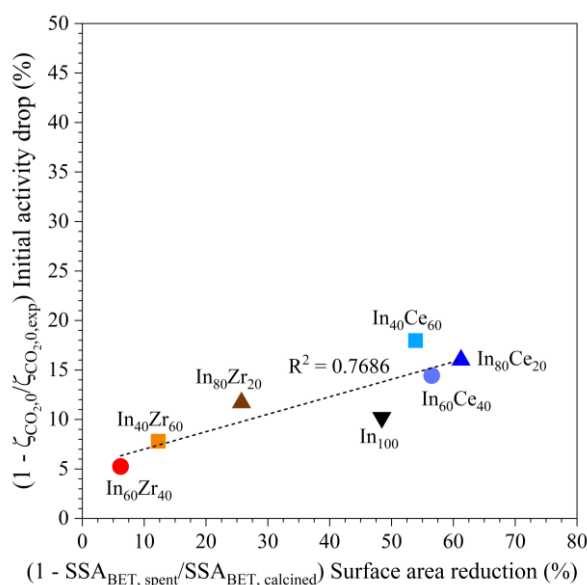


Figure 8. Relationship between the reduction of specific surface area during the catalytic tests under reaction conditions and the initial drop of activity during the stability test carried out at 275 °C. For the sake of clarity, $\zeta_{CO_2,0}$ was estimated by using equation (13); while $\zeta_{CO_2,0,exp}$ is the initial experimental CO₂ conversion.

423

424 As previously shown in Figure 3a, the spent In_xCe_{100-x} samples contain metallic indium. The semi-quantitative
 425 analysis of the X-ray diffractograms of the samples allowed to estimate the mass fraction of metallic indium
 426 compared to the total amount of indium and indium oxide in the samples. In more detail, the mass fraction of
 427 In₂O₃ reduced into metallic indium was 50%, 25%, 9% and 4% for the In₄₀Ce₆₀, In₆₀Ce₄₀, In₈₀Ce₂₀ and In₁₀₀

428 samples, respectively. Moreover, the H₂-TPR measurements revealed that the addition of CeO₂ increases the
 429 reducibility of the binary oxides at low temperatures (i.e., between 200 °C and 300 °C). Therefore, the
 430 relationship shown in Figure 9a demonstrates that the amount of metallic indium formed under reaction
 431 conditions depends directly on the reducibility at low temperature of the samples. This finding provides some
 432 tentative evidence that the reduction of In₂O₃ into metallic indium is a phenomenon slower than the loss of
 433 specific surface area and it is likely related to the gradual deactivation of the samples. Figure 9b displays indeed
 434 the relationship between the reducibility of the In_xCe_{100-x} samples and the deactivation constant at 275 °C.
 435 What is striking from those two correlations is that the reducibility at low temperature of the samples grows
 436 as the cerium content rises and promotes the formation of metallic indium on the surface of the samples, that
 437 leads to a continuous decreasing of the performances of the In_xCe_{100-x} catalysts. Concerning the In_xZr_{100-x}
 438 catalysts, the experimental evidence showed that ZrO₂ increases slightly the reducibility of the oxides at high
 439 temperatures, in fact, no metallic indium was detected in the spent catalysts. This result confirms that the
 440 deactivation of In_xZr_{100-x} samples is not linked to their reducibility.

441

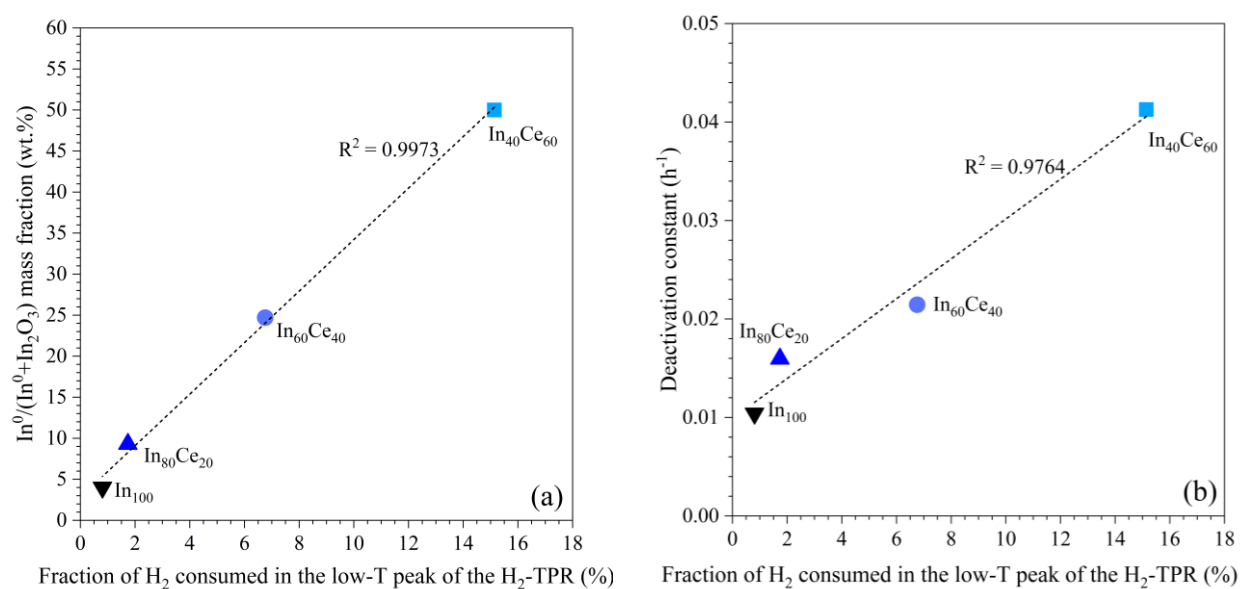


Figure 9. (a) Relationship between the fraction of H₂ consumed in the low-temperature peak of the H₂-TPR profiles and the mass fraction of metallic indium in the spent samples obtained by means of XRD. (b) Relationship between the fraction of H₂ consumed in the low-temperature peak of the H₂-TPR profiles and the deactivation constant at 275 °C for the In_xCe_{100-x} samples.

442

443 Another possible cause of deactivation could be coking. TPC is a suitable characterisation technique of the
444 spent catalysts to quantify the amount of carbon deposited on the surface of the samples during the catalytic
445 tests. Figure 10 illustrates the relationship between the amount of CO_x produced during the TPC and the
446 deactivation constant at 275 °C of the samples. What is interesting about the data in this figure is that the two
447 variables are clearly correlated to each other in the case of the $\text{In}_x\text{Ce}_{100-x}$ catalysts, whereas the deactivation
448 constant of $\text{In}_x\text{Zr}_{100-x}$ catalysts seems to be independent on the carbon deposition. More specifically, the
449 deactivation constant increases as the CeO_2 concentration in the sample grows.

450

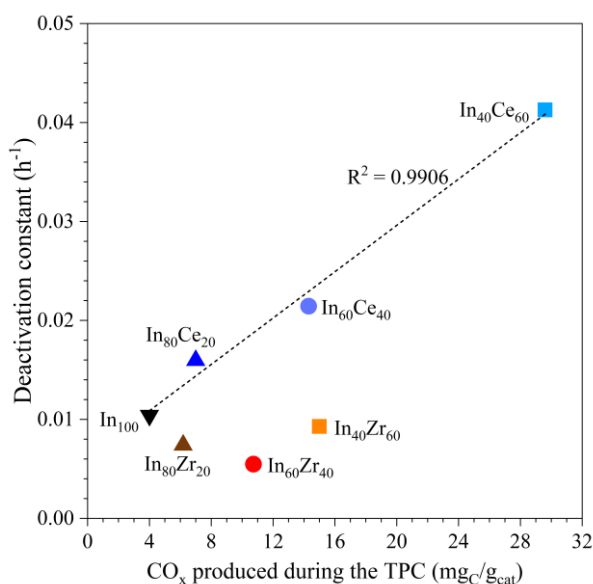


Figure 10. Relationship between the amount of CO and CO₂ produced during the TPC and the deactivation constant at 275 °C.

451

452 In summary, this section of the manuscript has attempted to provide an explanation to the deactivation
453 phenomena that occur on those catalysts during the CO₂ hydrogenation to methanol. It is worth noting that all
454 the samples seem to be severely affected by the reduction of the specific surface area during the initial part of
455 the test. Subsequently, $\text{In}_x\text{Zr}_{100-x}$ binary oxide catalysts reach steady performances at 275 °C with a low
456 deactivation constant and they did not seem to be affected by other deactivating phenomena. On the contrary,
457 the continuous decrease of the performances of the $\text{In}_x\text{Ce}_{100-x}$ binary oxide catalysts suggested that deactivation

458 phenomena slower than the structural shrinkage occurred. The reduction of In_2O_3 to metallic indium and the
459 carbon deposition were indeed the phenomena that almost certainly deactivated the samples.

460

461 3.3 Catalytic activity

462 As far as the catalytic activity tests are concerned, they were carried out at 2.5 MPa, $20 \text{ NL}\cdot\text{g}_{\text{cat}}^{-1}\cdot\text{h}^{-1}$ (inlet
463 composition: 60 vol.% H_2 , 20 vol.% CO_2 and 20 vol.% N_2), varying the reaction temperature between 200 °C
464 and 300 °C. Supplementary Figure S10 illustrates the catalytic performance with respect to the mass of catalyst
465 in the catalytic bed. From these results it is evident that the CO_2 conversion increases as the temperature rises
466 as a consequence of the increase of the reaction rate and it could be stated that the activity of the $\text{In}_x\text{Ce}_{100-x}$
467 samples decreases by reducing the indium content; whilst there is not a clear trend for the $\text{In}_x\text{Zr}_{100-x}$ samples.
468 To understand the effects of the addition of ZrO_2 or CeO_2 to In_2O_3 , Figure 11 displays the catalytic results with
469 respect to the actual mass of In_2O_3 in each sample.

470 Concerning the $\text{In}_x\text{Ce}_{100-x}$ samples, the specific CO_2 conversion rate in Figure 11a decreases proportionally as
471 the indium content diminishes, suggesting that the activity of In_2O_3 is worsened by the addition of CeO_2 . On
472 the contrary, $\text{In}_x\text{Zr}_{100-x}$ samples exhibited a specific CO_2 conversion rate greater than the one of the pure In_2O_3
473 sample. Moreover, the specific methanol productivity (see Figure 11b) is also enhanced by the addition of
474 ZrO_2 in the catalysts. Those results are consistent with the findings of Martin *et al.* [26]. The CO_2
475 hydrogenation to methanol is indeed boosted by the formation of oxygen vacancies on the surface of In_2O_3 ,
476 that is promoted by the presence of ZrO_2 [25,26]. The $\text{In}_{40}\text{Zr}_{60}$ catalyst reached a specific methanol productivity
477 of $168 \text{ mg}_{\text{CH}_3\text{OH}}\cdot\text{g}_{\text{In}_2\text{O}_3}^{-1}\cdot\text{h}^{-1}$ at 300 °C, while its methanol selectivity fell from 81.2 % at 200 °C to 28.3 % at
478 300 °C. However, as previously discussed, $\text{In}_x\text{Ce}_{100-x}$ samples were subjected to a significant reduction of the
479 specific surface area, while $\text{In}_x\text{Zr}_{100-x}$ catalysts kept a high specific surface area under reaction condition. Thus,
480 pure In_2O_3 showed the best specific methanol productivity per unit of surface area (i.e., $5.32 \text{ g}_{\text{CH}_3\text{OH}}\cdot\text{m}_{\text{cat}}^{-2}\cdot\text{h}^{-1}$
481 at 300 °C) followed by the $\text{In}_{80}\text{Ce}_{20}$ sample (i.e., $3.27 \text{ g}_{\text{CH}_3\text{OH}}\cdot\text{m}_{\text{cat}}^{-2}\cdot\text{h}^{-1}$ at 300 °C) and the $\text{In}_{80}\text{Zr}_{20}$ sample (i.e.,
482 $2.23 \text{ g}_{\text{CH}_3\text{OH}}\cdot\text{m}_{\text{cat}}^{-2}\cdot\text{h}^{-1}$ at 300 °C). Finally, it is worth mentioning that $\text{In}_x\text{Zr}_{100-x}$ catalysts produced small amounts
483 of dimethyl-ether as a by-product due to the surface acidity of ZrO_2 , which promotes the dehydration of

484 methanol to dimethyl-ether. The dimethyl-ether yield was 0.026%, 0.037% and 0.024% for the $\text{In}_{40}\text{Zr}_{60}$,
485 $\text{In}_{60}\text{Zr}_{40}$ and $\text{In}_{80}\text{Zr}_{20}$ catalysts, respectively.

486 Besides temperature, pressure is the other significant parameter of the methanol synthesis process. The latter
487 indeed occurs with a reduction of the number of moles as shown in equations (1) and (2), and according to the
488 Le Châtelier-Braun principle, methanol is a thermodynamically favoured product by the increase of pressure.
489 On the contrary, the RWGS reaction is not affected by the variation of the pressure. For these reasons, an
490 increase of the partial pressure of the reactants has a beneficial effect on the methanol selectivity; in fact,
491 Martin *et al.* [26] reported that the methanol selectivity approaches to 100% at 50 bar. Moreover, an increase
492 of the pressure promotes the reaction rates and increase the overall CO_2 conversion.

493

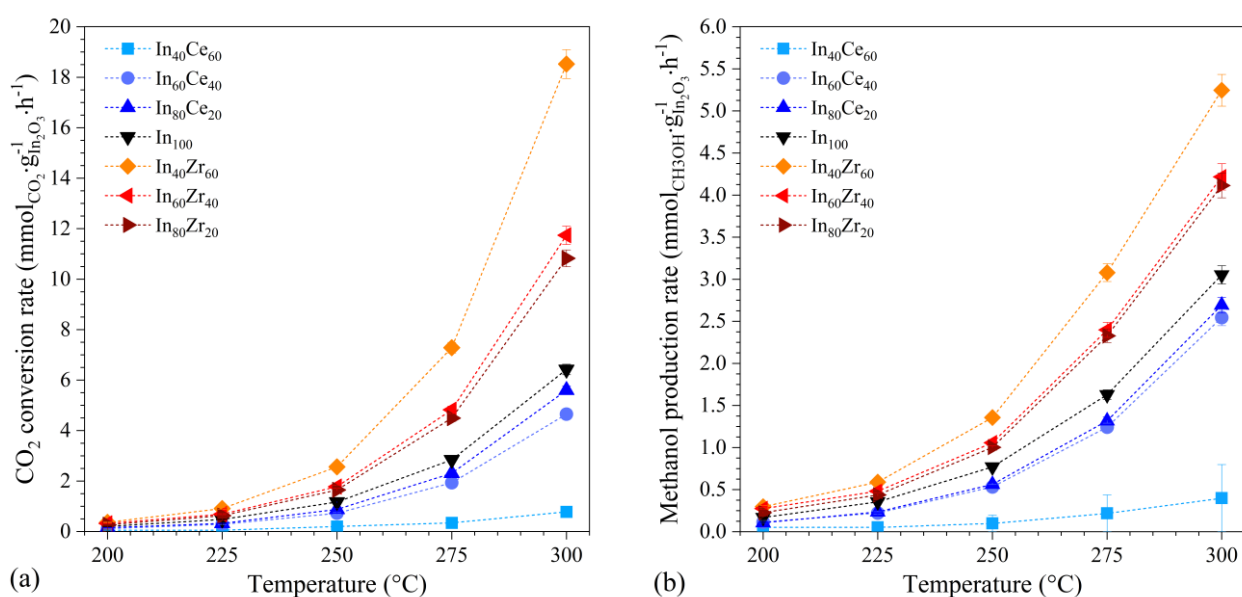


Figure 11. (a) Specific CO_2 conversion rate and (b) specific methanol production rate of the $\text{In}_x\text{Ce}_{100-x}$ and $\text{In}_x\text{Zr}_{100-x}$ binary oxide catalysts. Reaction conditions are 2.5 MPa and 20 NL·g $_{\text{cat}}^{-1}$ ·h $^{-1}$ (inlet composition: 60 vol.% H_2 , 20 vol.% CO_2 and 20 vol.% N_2). For the sake of clarity, the actual mass fractions of In_2O_3 estimated by means of ICP-MS measurements were used in the calculation of the specific CO_2 conversion rate.

494

495 As aforementioned, the performances of indium-based catalysts strongly depend on the presence of oxygen
496 vacancies; thus, to rationalise those intriguing findings related to the activity of the samples, we performed
497 XPS and Raman spectroscopy. The most significant results of those measurements are presented in Table 4,

498 while more detailed data are reported in Figure 2 and in the Supplementary Material (see Figure S5, Figure
499 S6, Figure S7 and Figure S8). Firstly, the superficial indium atomic ratio was estimated by using XPS and it
500 is worth noting that indium concentration on the surface of the $\text{In}_x\text{Ce}_{100-x}$ samples is higher than the one of the
501 $\text{In}_x\text{Zr}_{100-x}$ catalysts. In addition, if it is compared with bulk values obtained by means of ICP-MS measurements
502 (see Table 1), it is evident that the addition of CeO_2 increases the surface concentration of indium atoms. The
503 average oxidation state of indium estimated by using the modified Auger parameter is +3 in all the spent
504 samples. On the other hand, the average oxidation state of Ce estimated from the deconvolution of the Ce3d
505 signals of the $\text{In}_x\text{Ce}_{100-x}$ samples (see Figure S6a) decreases by rising the cerium content in the sample
506 demonstrating the increase in reducibility and the presence of Ce^{+3} . Whereas the average oxidation state of Zr
507 estimated from the deconvolution of the Zr3d signals in the $\text{In}_x\text{Zr}_{100-x}$ samples (see Figure S6b) is low,
508 suggesting the presence of Zr sub-oxides (ZrO_x with $x < 2$).

509 The concentration of oxygen defects was estimated from the deconvolution of the O1s signals of the XPS
510 measurements and from Raman spectroscopy. Concerning $\text{In}_x\text{Ce}_{100-x}$ samples, the deconvolution of the O1s
511 signal revealed that the concentration of oxygen defects increases from 14.0 at.% of the $\text{In}_{80}\text{Ce}_{20}$ to 21.6 at.%
512 of the $\text{In}_{40}\text{Ce}_{60}$ and this result is consistent with the Raman measurements. In more detail, the D/F_{2g} ratio of the
513 $\text{In}_x\text{Ce}_{100-x}$ samples is 0.03 for the calcined Ce_{100} powder, it is almost constant for the calcined binary oxides
514 (0.32 – 0.33) and increases up to 0.53 – 0.56 for all the spent catalysts. These data appear to support the idea
515 that oxygen defects are mainly linked to the cerium oxide rather than to the indium oxide. Therefore, the
516 addition of Ce in the samples increased the concentration of O defects in the CeO_2 structure and the reducing
517 atmosphere amplified the presence of O defects in the spent samples. Hence, the O defects in the $\text{In}_x\text{Ce}_{100-x}$
518 samples seems to be more correlated to the deactivation phenomena rather than the activity. In fact, if the
519 concentration of O defects increased in the CeO_2 phase while the In_2O_3 phase was only marginally affected by
520 this phenomenon, the activity in CO_2 hydrogenation would decrease as the indium content diminished.

521 On the other side, $\text{In}_x\text{Zr}_{100-x}$ samples appear to exhibit a different behaviour. Similar to the $\text{In}_x\text{Ce}_{100-x}$ samples,
522 the concentration of O defects in the spent $\text{In}_x\text{Zr}_{100-x}$ samples increases from 12.6 % for $\text{In}_{80}\text{Zr}_{20}$ to 21.1 % for
523 $\text{In}_{40}\text{Zr}_{60}$. However, the average oxidation state of Zr is significantly lower than the average oxidation state of
524 Ce. Hence, the O defects may also affect the In_2O_3 and not only the zirconium oxide. In support of this

525 interpretation, Martin *et al.* [26] have attributed the formation of O defects (and/or O vacancies) to the
 526 electronic promotion of Zr cations that change their oxidation state from +4 to +3, favouring the formation of
 527 other O defects in the In₂O₃ structure. Consequently, the increase of O defects in the indium oxide led to an
 528 increase of the catalytic activity in CO₂ hydrogenation.

529

530 Table 4. Results of XPS and Raman measurements on the spent samples.

Sample	In/(In+M) ^a	In Auger parameter ^b	M ⁴⁺ /M ^c	O _{defect} /O ^d	D/F _{2g} area ratio ^e
	at. %	eV	at. %	at. %	-
In ₄₀ Ce ₆₀	47.4	850.9	56.7	21.6	0.56
In ₆₀ Ce ₄₀	75.2	851.0	64.3	14.9	0.53
In ₈₀ Ce ₂₀	93.5	850.7	67.3	14.0	0.54
In ₁₀₀	100	850.7	-	17.1	-
In ₈₀ Zr ₂₀	80.0	851.2	15.2	12.6	-
In ₆₀ Zr ₄₀	70.2	850.4	30.4	17.5	-
In ₄₀ Zr ₆₀	44.0	850.7	21.6	21.1	-

^a Elemental atomic ratio on the surface of the samples estimated by means of XPS measurements. For the sake of clarity, M represents respectively Ce or Zr in the samples.

^b Modified Auger parameter of indium estimated by using the In MNN signals of the XPS measurements.

^c Fraction of Ce⁴⁺ or Zr⁴⁺ species with respect to Ce or Zr species estimated by means of the deconvolution of the Ce3d or Zr3d signals of the XPS measurements (see Figure S6).

^d Fraction of O_{defect} species with respect to O species estimated by means of the deconvolution of the O1s signal of the XPS measurements (see Figure S3).

^e Defects in the CeO₂ structure estimated by means of Raman spectroscopy (see Figure 2b).

531

532 To investigate if the reaction mechanism changes on the surface of the two catalytic groups, the experimental
 533 reaction rates were fitted on the Arrhenius' plot (see Figure S11, Figure S12 and Figure S13) and the results
 534 are reported in Table S3. What stands out from the table is that the apparent activation energies for the CO₂
 535 hydrogenation ranges between 66 kJ·mol⁻¹ and 89.9 kJ·mol⁻¹ with a relative uncertainty that achieves 15.2 %
 536 in the worst case (i.e., In₄₀Ce₆₀). Moreover, according to the literature [54], the apparent activation energy for
 537 CO synthesis is greater than the one for the methanol synthesis. In addition, it is worth noting that the pre-
 538 exponential factors and the activation energies are strongly dependent on each other; however, for each

539 reaction the apparent activation energy does not seem to change significantly. For these reasons, the pre-
540 exponential factors and the apparent activation energies were determined by assuming that the apparent
541 activation energy of each reaction is the same for all the samples and the results are summarised in Table 5.
542 What stands out from the table is that the apparent activation energy of the methanol synthesis ($65.5 \text{ kJ}\cdot\text{mol}^{-1}$)
543 is lower than the one of the RWGS reaction ($107.5 \text{ kJ}\cdot\text{mol}^{-1}$), while the apparent activation energy for the CO_2
544 hydrogenation is about $81.0 \text{ kJ}\cdot\text{mol}^{-1}$. A closer inspection of the pre-exponential factors revealed that the
545 addition of CeO_2 progressively reduces the number of active sites and consequently the performance of the
546 $\text{In}_x\text{Ce}_{100-x}$ samples. On the other hand, the pre-exponential factors of the $\text{In}_x\text{Zr}_{100-x}$ samples do not seem to
547 change at the variation of the ZrO_2 content; therefore, this feature resulted in a higher specific activity (i.e.,
548 number of active sites) of this group of catalysts. In conclusion, these findings are consistent with the previous
549 results of the characterisations and the higher performances of the $\text{In}_x\text{Zr}_{100-x}$ samples could be almost certainly
550 ascribed to the electronic promotion of Zr.

551

552 Table 5. Pre-exponential factors ($k_{\infty,i,j}^*$) and apparent activation energies ($E_{A,i}$) estimated by means of the
553 minimization of the objective function presented in equation (16) by determining a unique value of apparent
554 activation energy for each reaction.

Catalyst	CO_2 hydrogenation			CO synthesis			Methanol synthesis		
	E_{A,CO_2}	$\ln(k_{\infty,\text{CO}_2,j}^*)^a$	R^2	$E_{A,\text{CO}}$	$\ln(k_{\infty,\text{CO},j}^*)^a$	R^2	$E_{A,\text{CH}_3\text{OH}}$	$\ln(k_{\infty,\text{CH}_3\text{OH},j}^*)^a$	R^2
	$\text{kJ}\cdot\text{mol}^{-1}$	-	-	$\text{kJ}\cdot\text{mol}^{-1}$	-	-	$\text{kJ}\cdot\text{mol}^{-1}$	-	-
$\text{In}_{40}\text{Ce}_{60}$	81.0 ± 0.2	22.77 ± 0.36	0.8861	107.5 ± 0.2	27.68 ± 0.40	0.5169	65.5 ± 0.2	18.82 ± 0.34	0.8251
$\text{In}_{60}\text{Ce}_{40}$		24.71 ± 0.09	0.9957		29.51 ± 0.08	0.9958		20.81 ± 0.13	0.9871
$\text{In}_{80}\text{Ce}_{20}$		25.30 ± 0.12	0.9921		30.39 ± 0.26	0.9717		21.24 ± 0.14	0.9854
In_{100}		25.76 ± 0.07	0.9963		30.73 ± 0.07	0.9981		21.74 ± 0.04	0.9986
$\text{In}_{80}\text{Zr}_{20}$		26.03 ± 0.11	0.9934		31.07 ± 0.10	0.9969		21.92 ± 0.08	0.9939
$\text{In}_{60}\text{Zr}_{40}$		25.94 ± 0.12	0.9913		30.96 ± 0.08	0.9980		21.81 ± 0.10	0.9904
$\text{In}_{40}\text{Zr}_{60}$		25.91 ± 0.16	0.9868		31.09 ± 0.15	0.9929		21.65 ± 0.07	0.9954

^a The unit of the pre-exponential factor ($k_{\infty,i,j}^*$) are $\text{mmol}\cdot\text{kg}_{\text{cat}}^{-1}\cdot\text{h}^{-1}$.

555

556 4 Conclusions

557 This study has examined the differences between two groups of In_2O_3 -based binary oxides catalysts to convert
558 CO_2 and H_2 into methanol. The catalysts were tested in a fixed-bed reactor to evaluate their performances in
559 terms of activity, selectivity and stability. Then, all the samples were analysed with several characterisation
560 techniques to rationalise the experimental results. First of all, the addition of Ce and Zr in the catalysts modifies
561 completely their morphology and textural properties. XRD measurements have suggested the formation of an
562 amorphous In_2O_3 - ZrO_2 solid solution for the $\text{In}_x\text{Zr}_{100-x}$ samples, while they have revealed the formation of an
563 In-doped CeO_2 phase in the other group of catalysts, that was also attested by Raman spectroscopy.

564 Moreover, the adsorption of CO_2 was amplified by the addition of Ce or Zr from $1.8 \text{ mmol}_{\text{CO}_2} \cdot \text{g}_{\text{cat}}^{-1}$ of the pure
565 In_{100} up to $10.6 \text{ mmol}_{\text{CO}_2} \cdot \text{g}_{\text{cat}}^{-1}$ of the $\text{In}_{40}\text{Ce}_{60}$ or $6.6 \text{ mmol}_{\text{CO}_2} \cdot \text{g}_{\text{cat}}^{-1}$ of the $\text{In}_{40}\text{Zr}_{60}$; however, this fact results in
566 a different behaviour of the two groups of catalysts. This feature has been ascribed to the higher concentration
567 of oxygen defects and/or vacancies, that have been observed by using XPS and FT-IR measurements and
568 Raman spectroscopy and supported by the literature. In more detail, the results have supported the idea that Zr
569 has an electronic promotion effect on the In_2O_3 by increasing its specific activity. Moreover, the addition of Zr
570 in the catalysts significantly increases the specific surface area and stabilizes the structure of the material,
571 mitigating the effects of sintering. As a result, the deactivation phenomena have been attributed to the loss of
572 specific surface area. On the other hand, the experimental evidence has revealed that $\text{In}_x\text{Ce}_{100-x}$ samples were
573 affected by at least three deactivation mechanisms: sintering of the structure, formation of metallic indium and
574 coking. More specifically, the deactivation rate increases as the amount of Ce rises; besides, the addition of Ce
575 in the catalysts increased their reducibility under reaction condition promoting the reduction of In_2O_3 to
576 metallic indium, that was detected by XRD measurements.

577 The insights gained from this study may be of assistance to the development of more stable and active catalysts
578 for the CO_2 conversion into methanol. Moreover, the beneficial effect of ZrO_2 could expand the operating
579 condition range for methanol synthesis to develop more efficient processes of one-step processes to produce
580 other added-value products from CO_2 and H_2 .

581 **Acknowledgements**

582 The authors would like to acknowledge the Italian regional project SATURNO – Bioeconomy (Piedmont
583 Region 2019-2022, <https://saturnobioeconomia.it/>) and the project PON Ricerca e Innovazione "REACT-EU"
584 (DM 1062/21) funded by the Italian Ministero dell'Università e della Ricerca (MUR).

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