## **Electronic Supplementary Information**

for

## Organic bases catalyze the synthesis of urea from ammonium salts derived from recovered environmental ammonia

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	0 NH4 <sup>+ -</sup> 0 <sup></sup> NH <sub>2</sub>	Catalyst O NMP, 140 °C, 24 h H <sub>2</sub> N	NH <sub>2</sub> + H <sub>2</sub> O
Entry	Base	p <i>K</i> <sub>a</sub> values of conjugated acid	Yield (%)
1	DBU	24.3	18
2	TMG	23.4	4
3	TBD	25.5	3
4	MTBD	26.0	3
5	DBN	23.4	3
6	$Et_3N$	18.8	Trace
7	DMAP	18.0	Trace
8	Proton sponge	N/A	Trace
9	DABCO	N/A	Trace
10	-	-	Trace

**Figure S1.** Urea synthesis from ammonium carbamate catalyzed by organic bases. Experimental conditions: Base (0.38 mmol), ammonium carbamate (3.8 mmol), 140 °C, 24 h, in 1 mL of NMP. Pressure of the inside of the vessel was increased up to 0.48 MPa because of autogenous pressure of thermal decomposition of ammonium carbamate. The amount of produced area was determined by the Fearon reaction.

		MP, 140 °C, 24 h H₂N NI	H <sub>2</sub> + H <sub>2</sub> O
Entry	Solvent	Relative dielectric constant	Yield (%)
1	-	-	0
2	DMSO	46.5	1
3	MeCN	35.9	0
4	NMP	32.2	0
5	THF	7.6	0
6	Toluene	2.4	0
10	1,4-dioxane	2.2	0

0

Figure S2. Urea synthesis from ammonium carbamate without base.

0

Experimental conditions: Ammonium carbamate (3.6 mmol), 140 °C, 5 atm, 24 h, in 1 mL of solvent.



Figure S3. Substrate Scope.



**Figure S4.** <sup>13</sup>C NMR spectra in DMSO of (A)  $Et_3N + H_2SO_4$ , (B)  $Et_3N + ammonium$  carbamate, and (C)  $Et_3N$ . Filled circle: primary atom of  $Et_3NH^+$ , Asterisk: primary carbon atom of  $Et_3N$ . The signal of the carbon of protonated  $Et_3N$  is observed at around 8.7 ppm.



**Figure S5.** FT-IR spectra of (A)  $Et_3N$ , (B) ammonium carbamate, and (C)  $Et_3N$  + ammonium carbamate. Dashed line: C=O stretch of carbamate anion.



**Figure S6.** Urea synthesis results under various temperature as a function of time. Experimental conditions: DBU (0.36 mol), ammonium carbamate (3.6 mol), 5 atm, in 1 mL of DMSO. Black (140 °C), Green (130 °C), Red (120 °C), Blue (110 °C).

Entry	Reaction temperature	Slope (%/h)	<i>v</i> <sub>0</sub> (mol/(L*h))	ln( <i>v</i> <sub>0</sub> )
1	110 °C	1.96	0.775	-2.56
2	120 °C	3.70	0.142	-1.95
3	130 °C	5.30	0.198	-1.62
4	140 °C	7.47	0.272	-1.30

Figure S7. Urea synthesis results under various temperature.